Chemistry 115 Name

Dr. Cary Willard

Quiz 8a (20 points) November 21, 2012

1. (5 points) Calculate the mass percent for the solute if 35.0 g of KCl are added to 375 g of water.
2. (5 points) Calculate the molarity of a solution prepared by dissolving 74.3 grams of glucose (C6H12O6) in enough wáter to make 3.59 L of solution.
3. (5 points) Calculate the mass of sodium nitrate required to prepare 425 mL of a 0.377 M NaNO3 solution.
4. (5 points) A 45.0 mL sample of 3.92 M potassium sulfate is diluted to a final volume of 626 mL. What is the molarity of K2SO4 in the final solution?

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Quiz 8b (20 points) November 21, 2012

1. (5 points) Calculate the mass percent for the solute if 65.0 g of KCl are added to 875 g of water.
2. (5 points) Calculate the molarity of a solution prepared by dissolving 67.4 grams of glucose (C6H12O6) in enough wáter to make 2.88 L of solution.
3. (5 points) Calculate the mass of sodium nitrate required to prepare 725 mL of a 0.377 M NaNO3 solution.
4. (5 points) A 75.0 mL sample of 3.92 M potassium sulfate is diluted to a final volume of 626 mL. What is the molarity of K2SO4 in the final solution?

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Quiz 8c (20 points) November 21, 2012

1. (5 points) Calculate the mass percent for the solute if 22.8 g of AgNO3 are added to 942 g of water.
2. (5 points) Calculate the molarity of a solution prepared by dissolving 83.4 grams of fructose (C6H12O6) in enough wáter to make 5.77 L of solution.
3. (5 points) Calculate the volume (mL) of 1.37 M NaNO3 solution that will contain 82.5 g of sodium nitrate.
4. (5 points) To what volume must 25.0 mL of a 8.15 M solution potassium sulfate be diluted to make a solution with a final concentration of 2.85 M?

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Quiz 8d (20 points) November 21, 2012

1. (5 points) Calculate the mass percent for the solute if 38.4 g of AgNO3 are added to 942 g of water.
2. (5 points) Calculate the molarity of a solution prepared by dissolving 97.3 grams of fructose (C6H12O6) in enough wáter to make 5.77 L of solution.
3. (5 points) Calculate the volume (mL) of 2.16 M NaNO3 solution that will contain 82.5 g of sodium nitrate.
4. (5 points) To what volume must 25.0 mL of a 9.15 M solution potassium sulfate be diluted to make a solution with a final concentration of 2.85 M?